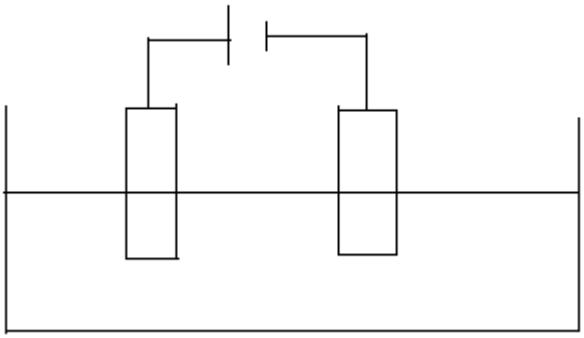
ELECTROLYSIS:

A substance's particles are moved apart from each other hence a electrical current is applicable

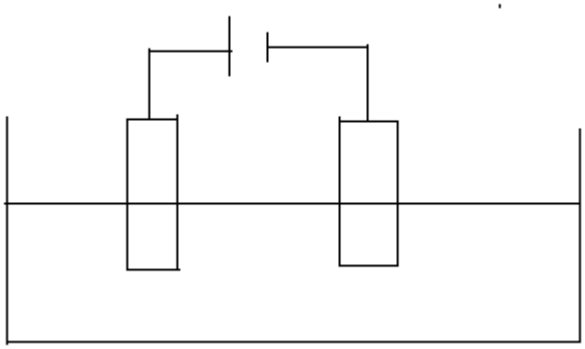


The higher the concentration of the ions the greater the current

- All ionic substances
- All acids

Types of reactions:

1) **Electrolysis**
   - The breakdown of a substance's and its components through an electrical current.



e.g. $CuCl_2 \longrightarrow Cu + Cl_2$

2) **Precipitation Reaction:**

Test for Cl^-, Br^-, and I^- - $AgNO_3$ is added.
Test for SO_4^{2-} en CO_3^{2-} - $BaCl_2$ is added
Precipitate is insoluble.

$$ZnCl_{2aq} + AgNO_{3aq} \longrightarrow Zn(NO_3)_2 + 2AgCl$$

$$MgSO_{4aq} + BaCl_{2aq} \longrightarrow MgCl_{2aq} + BaSO_4$$

3) <u>Gas:</u>

a) Metal + acid -> salt + water

b) $$\text{Metaal-karbonaat} + \text{Suur} \longrightarrow \text{Sout} + H_2O + CO_2$$

Test for O_2: Glowing splinter
Test H_2: A wooden splinter catches light
Tests CO_2: Lime water turns milky

4) <u>Acid - Base:</u>

Acid: Starts with H.
Base: Ends on OH.

Eg. $HCl + NaOH \longrightarrow NaCl + H_2O$

$\text{Suur} + \text{Basis} \longrightarrow \text{Sout} + \text{Water}$

5) <u>Redox Reaction:</u>

Charges change of a element
Eg. $N_2 + 3H_2 \longrightarrow 2NH_3$